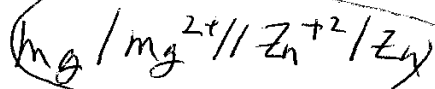


Electrochemistry lab



Galvanic Electrochemical Cell

cell #1 measure volt

$$E^{\circ}_{\text{cell}} = 2.37 + (-0.763)$$

$E^{\circ}_{\text{cell}} = +1.627$
expt
reaction should go forward

negative electrode
(-) black wire

oxidation
(expt setup)
use cup

metal	solution
Mg	$\text{Mg}(\text{NO}_3)_2$ (aq)

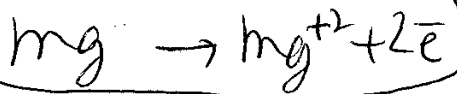
positive electrode

(+) red wire

reduction
(expt. setup)
use beaker

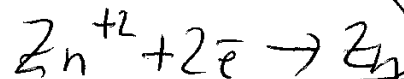
metal	solution
Zn	ZnSO_4 (aq)

oxidation

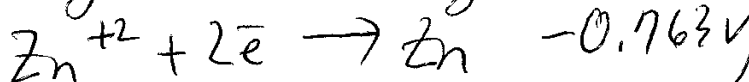
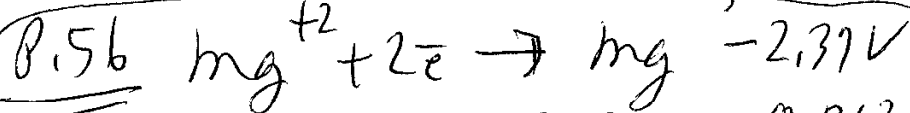


+2.37

reduction



-0.763



from chart