Name Ker	~	Name
(print name)	1	(sign name) Please show all work for full (& partial) cred
1. Is the follow	ing an allowed Lewis Dot Str	ucture? Pay attention to the atom with the *. Gellow
_4 1 / 2	inala anal (5 nta)	ed octet (b) less than an octet structure (c) incorrect
	e expande	a elevent - con
2. Place the fol pts) (a) lo	lowing bonds in order using tongest bond length (b) mediur	he choices below (each letter can only be used one time) (an bond length (c) shortest bond length
C=C		n bond length (c) shortest bond length
c-c_a		(no pairs)
C=C _C		love
S	Counts as 1	or the Lewis Dot structure shown. Sine dash wedge Vector + 0 Sum.
	ng questions about the atom v	ţ
	1	b. How many lone pairs
c. What is the	geometry of the electron pairs	or trisonal planar
d. What is the	geometry of the molecule?	bert
e. What is the	hybridization <u>Sp</u> 2	f. What is, are the bond angle $\frac{120^{\circ}}{}$
Extra Credit (1 pt e	ach, 3 pts) For the bond, sho	w the polarity symbol as an arrow or with the symbol δ
SO		

Name		en		Name		*******		
(print n	ame)	0		(sign nar	ne) Pleas	e show all v	vork for full (& partial)	cred
1.	Is the follo	wing an all	owed Lewis Do	ot Structure?	Pay atter	ation to the	atom with the *.	
		Dot structi (circle one)		panded octet (b) less tha	an an octet s	tructure(c) incorrect	
ē	6 f	λ	I campo	t expc	ed a	estet		
F —	N*	2	ind pe	riod e	rlene	ist		
,	F.					(Apr		
2.	Place the f pts) (a	ollowing be a) longest b	onds in order us	sing the choice medium bond	es below (length (c)	each letter of shortest bo	an only be used one tine	ne) (
	c-c @	<u> </u>	C=C _(C=C	<i>(</i> 2)			
3.	Complete	the followi	ng VSEPRT ch	(201/201	ia		vn. (ecto-+0) sun	N Ve - 8
11	* H			one pair	H	H	As in	4 <i>l</i> +
Answe	r the follov	ving questi	ons about the at	om with the *	(2 pts eac	ch, 12 pts)		/
a.	How many	y electron p	pairs 4	b.	How ma	any lone pai		
c.	What is th	e geometry	of the electron	pairs ?	etve	ebelly	ol	
d.	What is th	e geometry	of the molecul	e? tvi	Gon	al p	y ramidal ngle 109,5°	
e.	What is th	e hybridiza	tion $S\rho^3$	f. W	'hat is, are	the bond ar	$\frac{109.5^{\circ}}{}$	_
							row or with the symbol	δ
H	As							
	7							
+	,		or the molecule					

Name	Name	
(print na	hell.	
1.	s the following an allowed Lewis Dot Structure? Pay attention to the atom with the *.	
	The Lewis Dot structure is an [(a) expanded octet (b) less than an octet structure (c) incorrect tructure] (circle one) (5 pts)	
	group IIA is often less	
	group IIA is often less than order	
H/	H H	
11,	· ·	
	Place the following bonds in order using the choices below (each letter can only be used one time) (3	
	ots) (a) Highest bond strength (b) medium bond strength (c) lowest bond strength	
ı	$C=C \bigcirc \qquad \qquad C=C \bigcirc \qquad \qquad C=C \bigcirc \qquad \qquad \qquad C=C \bigcirc \qquad \qquad$	
3.	Complete the following VSEPRT chart for the Lewis Dot structure shown.	
	Br. (no vector) Br.	
*	Co-	
	Br. Br. Br. Br. Br. Br. Br. Br.	
Br	Br Br	
-	Br	
Answer	the following questions about the atom with the * (2 pts each, 12 pts)	
a.	How many electron pairs b. How many lone pairs	
c.	What is the geometry of the electron pairs? <u>Frigoral Dipyramidal</u>	
	What is the geometry of the molecule?	
e.	What is the hybridization $\frac{50^3 d}{10^3 d}$ f. What is, are the bond angle $\frac{20^3 + 90^5}{10^3}$	
Extra C	redit (1 pt each, 3 pts) For the bond, show the polarity symbol as an arrow or with the symbol δ	
Br	Cl	
	`	
+		

Quiz V	II General Chemistry I Lecture Fall 13 Dr. Hahn 20 pts 11/25 M 10:30 am Form B quiz #
Name	K & Name
(print r	
· ·	Is the following an allowed Lewis Dot Structure? Pay attention to the atom with the *.
	The Lewis Dot structure is an [(a) expanded octet (b) less than an octet structure (c) incorrect structure] (circle one) (5 pts)
H-/	will accept either @ or @ We ally O- too few & struture + 0 not other
2.	Place the following bonds in order using the choices below (each letter can only be used one time) (3 pts) (a) Highest bond strength (b) medium bond strength (c) lowest bond strength
	$C=C \bigcirc \qquad \qquad C=C \qquad C=C $
3.	Complete the following VSEPRT chart for the Lewis Dot structure shown.
Cl	Cl: $Vector Sum = 0$ Cl: Cl Cl:
Answe	er the following questions about the atom with the * (2 pts each, 12 pts)
a.	How many electron pairs b. How many lone pairs
c.	What is the geometry of the electron pairs? Octabed val What is the geometry of the molecule? Octabed val
d.	What is the geometry of the molecule?
e.	What is the hybridization $\frac{5p^3d^2}{}$ f. What is, are the bond angle $\frac{900}{}$
Extra	Credit (1 pt each, 3 pts) For the bond, show the polarity symbol as an arrow or with the symbol δ
Cl	
	in the vector arrows for the molecule above in the line- wedge- dash drawing.
The ab	structure, vectors sum=0

Name	Name
(print name)	(sign name) Please show all work for full (& partial) credit
1. Is the following an all	lowed Lewis Dot Structure? Pay attention to the atom with the *. " Hellow
The Lewis Dot structustructure] (circle one)	ure is an [(a) expanded octet (b) less than an octet structure (c) incorrect) (5 pts)
Cl.	
CI CI	
ci :	
Cl	
2 Place the following by	onds in order using the choices below (each letter can only be used one time) (3
	nd length (b) medium bond length (c) shortest bond length
C=C	
C-C	
C≡C	
3 Complete the following	ng VSEPRT chart for the Lewis Dot structure shown.
e ¢	S Cline dash wedge
*	
0, 0,	0 0
Answer the following question	ons about the atom with the * (2 pts each, 12 pts)
a. How many electron p	pairs b. How many lone pairs
c. What is the geometry	of the electron pairs?
d. What is the geometry	of the molecule ?
e. What is the hybridiza	f. What is, are the bond angle
Extra Credit (1 pt each, 3 pts	For the bond, show the polarity symbol as an arrow or with the symbol δ

me	Name			
int name)	(sign	name) Ple	ase show all work	for full (& partial) credit
1. Is the following an allowed	Lewis Dot Structure	? Pay atte	ention to the atom	with the *.
The Lewis Dot structure is a structure] (circle one) (5 pt		et (b) less t	han an octet struct	ure (c) incorrect
F			•	
F				
2. Place the following bonds in pts) (a) longest bond le	order using the chongth (b) medium bo	oices below nd length ((each letter can o c) shortest bond le	nly be used one time) (3 ength
C-C C=C	C=	C		
3. Complete the following VS	EPRT chart for the I	Lewis Dot	structure shown.	
			**	
H H			As Maring H	
HAsH		н 🔨	Marian .	
H H		п	Н	
swer the following questions ab	out the atom with the	e * (2 pts e	ach, 12 pts)	
a. How many electron pairs _	b.	How 1	nany lone pairs	
c. What is the geometry of the	electron pairs ?			
d. What is the geometry of the	molecule ?	.4		<u>.</u>
e. What is the hybridization _	f.	What is, a	re the bond angle	
tra Credit (1 pt each, 3 pts) For	the bond, show the	polarity sy	mbol as an arrow	or with the symbol δ
As				
aw in the vector arrows for the r	olecule above in th	e line- wed	ge- dash drawing.	

The above molecule as a whole is [(polar) or (nonpolar)] (circle one)

Quiz VII General Chemistry I Lecture Fall I	3 Dr. Hahn 20 pts 11/25 M 10:30 am Form B quiz #
Name	Name
(print name)	Name (sign name) Please show all work for full (& partial) credi
1. Is the following an allowed Lewis Dot	Structure? Pay attention to the atom with the *.
The Lewis Dot structure is an [(a) expastructure] (circle one) (5 pts)	anded octet (b) less than an octet structure (c) incorrect
H O*	
	ng the choices below (each letter can only be used one time) (3 edium bond strength (c) lowest bond strength
C=C C-C	C=C
3. Complete the following VSEPRT char	t for the Lewis Dot structure shown.
CI CI CI CI CI	
Answer the following questions about the ator	n with the * (2 pts each, 12 pts)
a. How many electron pairs	b. How many lone pairs
c. What is the geometry of the electron pa	airs?
d. What is the geometry of the molecule	?
e. What is the hybridization	f. What is, are the bond angle
Extra Credit (1 pt each, 3 pts) For the bond, s	show the polarity symbol as an arrow or with the symbol δ
C1S	
Draw in the vector arrows for the molecule ab	ove in the line- wedge- dash drawing.
The above molecule as a whole is [(polar) o	or (nonpolar)] (circle one)