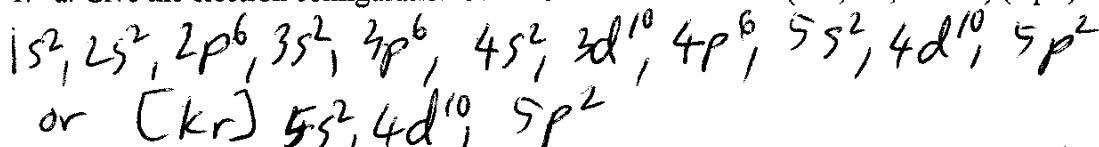


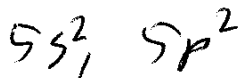
Birange

Name Key (print name) Name _____ (sign name) (I can't read some of your handwriting.)

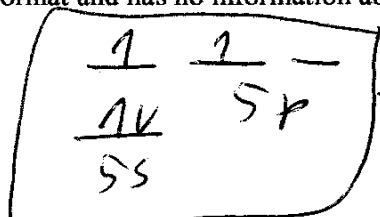
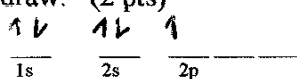
1. a. Give the electron configuration of Sn in the format (1s², 2s², ... etc.) (2 pts)



b. Give the valence electron configuration of the same element in the same format. (2 pts)



c. Give the valence electron configuration diagram for the same element in the format shown. The drawing below is only to show format and has no information about the correct structure which you need to draw. (2 pts)



ok if drew 5s + 5p in same level -

2. Periodic Relationships: Circle the one correct choice. (4 pts)

(As) or (N)] has the larger atom size (circle one)

(O) or (Se)] has the higher electronegativity (circle one)

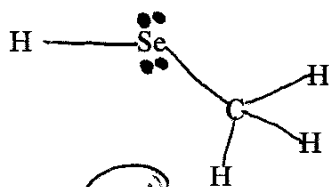
3. Give the Lewis Dot Symbol for the element $\cdot\overset{\cdot\cdot}{\underset{\cdot\cdot}{P}}\cdot$ (4 pts)

P in group 5 - has 5 dots

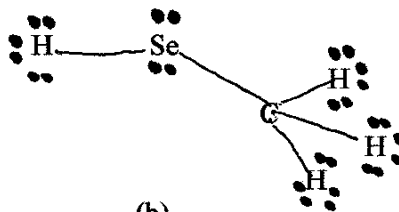
4. a. For the formula HSeCH₃ give the number of total valence electrons. Show work. (3 pts)

1 + 6 + 4 + 3(1) = 14e⁻ H group # 1
H Se C ↑ Se group # 6
C group # 4

b. Given the two structures drawn for the formula given above, circle the letter under the correct Lewis Dot structure (3 pts)



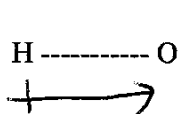
7x2 = 14e⁻



H can only have one lone pair

19x2 = 38e⁻ - too many e⁻

Extra Credit: Give the dipole moment arrow (+--->) for the bond shown below. (3 pts)

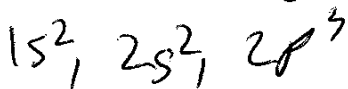


O is more EN than H

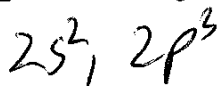
H is more EN as C

Name Key Name _____
 (print name) (sign name) (I can't read some of your handwriting.)

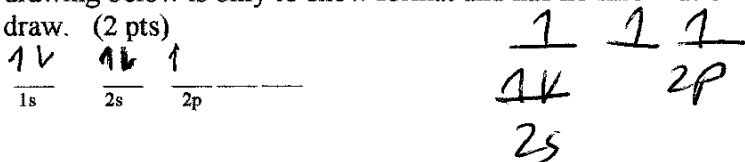
1. a. Give the electron configuration of N in the format (1s², 2s², ... etc.) (2 pts)



b. Give the valence electron configuration of the same element in the same format. (2 pts)



c. Give the valence electron configuration diagram for the same element in the format shown. The drawing below is only to show format and has no information about the correct structure which you need to draw. (2 pts)



2. Periodic Relationships: Circle the one correct choice. (4 pts)

[B] or (N) has the smaller atom size (circle one)

[(Al)] or (Cl) has the lower electronegativity (circle one)

3. Give the Lewis Dot Symbol for the element Br (4 pts)

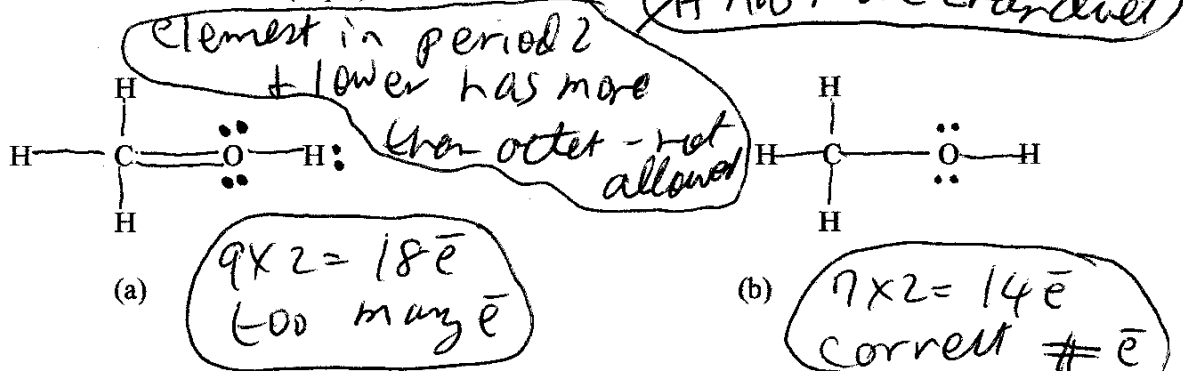
Br in group 7 - has 7 dots

4. a. For the formula H₃COH give the number of total valence electrons. Show work. (3 pts)

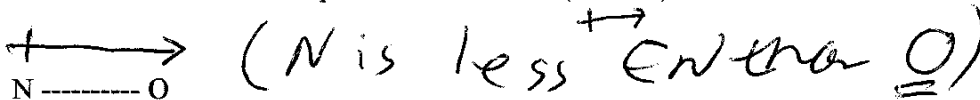
$$3(1e) + 4e + 6e + 1e = 14e$$

H C O H

b. Given the two structures drawn for the formula given above, circle the letter under the correct Lewis Dot structure (3 pts)



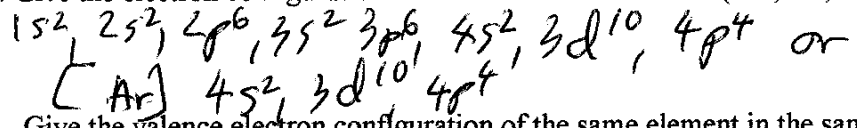
Extra Credit: Give the dipole moment arrow (+---->) for the bond shown below. (3 pts)



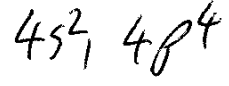
gold

Name key Name _____
(print name) (sign name) (I can't read some of your handwriting.)

1. a. Give the electron configuration of Se in the format (1s², 2s², ... etc.) (2 pts)



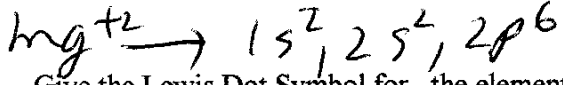
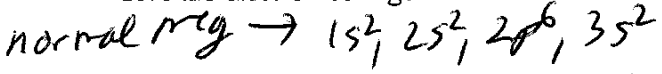
b. Give the valence electron configuration of the same element in the same format. (2 pts)



c. Give the valence electron configuration diagram for the same element in the format shown. The drawing below is only to show format and has no information about the correct structure which you need to draw. (2 pts)



2. Give the electron configuration for the ion Mg²⁺ (in the format 1s², 2s², 2p²...) (4 pts)

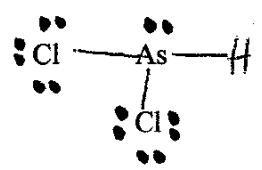
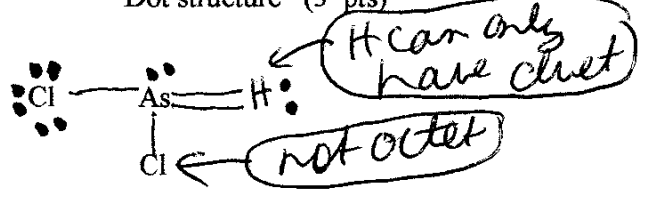


3. Give the Lewis Dot Symbol for the element • Mg • (4 pts)

4. a. For the formula AsCl₂H give the number of total valence electrons. Show work. (3 pts)

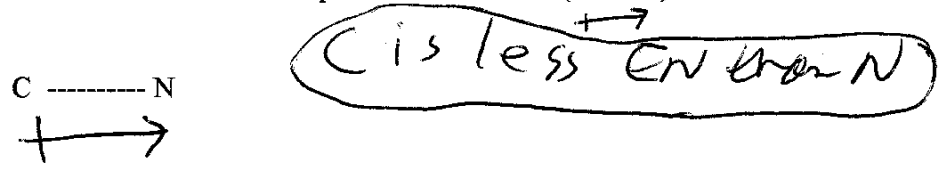
5e + (7e)2 + 1e = 20e
As Cl H

b. Given the two structures drawn for the formula given above, circle the letter under the correct Lewis Dot structure (3 pts)



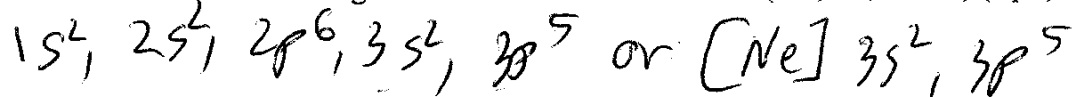
(b) $10 \times 2 = 20e$

Extra Credit: Give the dipole moment arrow (+---->) for the bond shown below. (3 pts)

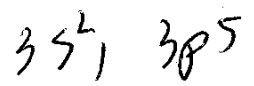


Name Key Name _____
 (print name) (sign name) (I can't read some of your handwriting.)

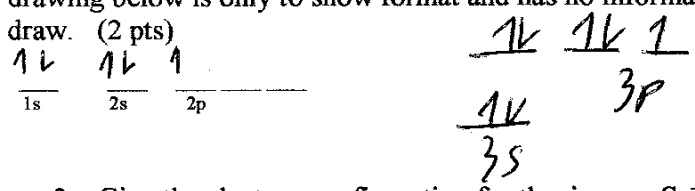
1. a. Give the electron configuration of **Cl** in the format (1s², 2s², ... etc.) (2 pts)



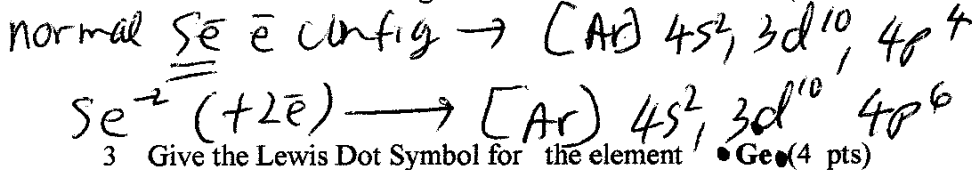
b. Give the valence electron configuration of the same element in the same format. (2 pts)



c. Give the valence electron configuration diagram for the same element in the format shown. The drawing below is only to show format and has no information about the correct structure which you need to draw. (2 pts)

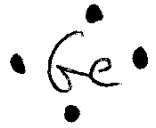


2. Give the electron configuration for the ion **Se²⁻** (in the format 1s², 2s², 2p²....) (4 pts)



3. Give the Lewis Dot Symbol for the element **Ge** (4 pts)

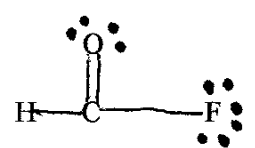
Ge in group 4
 ↳ has 4 dots



4a. For the formula **HCOF** give the number of total valence electrons. Show work. (3 pts)

(1e⁻) + (4e⁻) + 6e⁻ + 7e⁻ = 18e⁻
 H C O F

b. Given the two structures drawn for the formula given above, circle the letter under the correct Lewis Dot structure (3 pts)

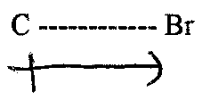


period 2 + lower element has more than octet
 O has 10e⁻ + C has 10e⁻
 not allowed

(a) 9 × 2 = 18e⁻

(b) 11 × 2 = 22e⁻

Extra Credit: Give the dipole moment arrow (+--->) for the bond shown below. (3 pts)



C is less EN than Br

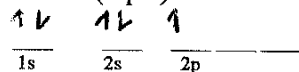
Birangl

Name _____ Name _____
 (print name) (sign name) (I can't read some of your handwriting.)

1. a. Give the electron configuration of **Sn** in the format ($1s^2, 2s^2, \dots$ etc.) (2 pts)

- b. Give the **valence** electron configuration of the same element in the same format. (2 pts)

- c. Give the **valence** electron configuration diagram for the same element in the format shown. The drawing below is only to show format and has no information about the correct structure which you need to draw. (2 pts)



2. Periodic Relationships: Circle the one correct choice. (4 pts)

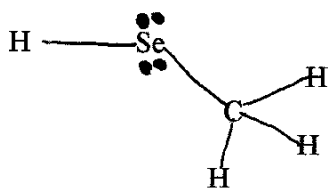
[(As) or (N)] has the larger atom size *(circle one)*

[(O) or (Se)] has the higher electronegativity *(circle one)*

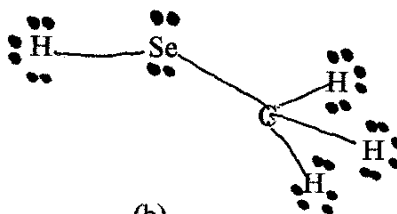
3. Give the Lewis Dot Symbol for the element **P** (4 pts)

4. a. For the formula **HSeCH₃** give the number of total valence electrons. Show work. (3 pts)

- b. Given the two structures drawn for the formula given above, circle the letter under the correct Lewis Dot structure (3 pts)

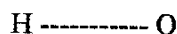


(a)



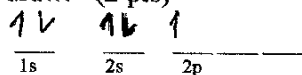
(b)

Extra Credit: Give the dipole moment arrow ($\text{+} \text{---} \text{>}$) for the bond shown below. (3 pts)



Name _____ Name _____
 (print name) (sign name) (I can't read some of your handwriting.)

1. a. Give the electron configuration of N in the format (1s², 2s², ... etc.) (2 pts)
- b. Give the valence electron configuration of the same element in the same format. (2 pts)
- c. Give the valence electron configuration diagram for the same element in the format shown. The drawing below is only to show format and has no information about the correct structure which you need to draw. (2 pts)



2. Periodic Relationships: Circle the one correct choice. (4 pts)

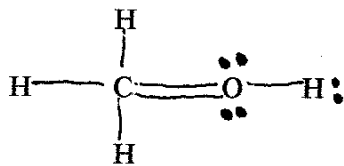
[(B) or (N)] has the smaller atom size (circle one)

[(Al) or (Cl)] has the lower electronegativity (circle one)

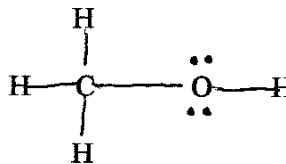
3. Give the Lewis Dot Symbol for the element Br (4 pts)

4. a. For the formula H₃COH give the number of total valence electrons. Show work. (3 pts)

- b. Given the two structures drawn for the formula given above, circle the letter under the correct Lewis Dot structure (3 pts)



(a)



(b)

Extra Credit: Give the dipole moment arrow (+---->) for the bond shown below. (3 pts)

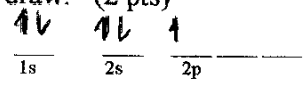


Name _____ Name _____
 (print name) (sign name) (I can't read some of your handwriting.)

1. a. Give the electron configuration of **Se** in the format (1s², 2s², ... etc.) (2 pts)

- b. Give the valence electron configuration of the same element in the same format. (2 pts)

- c. Give the valence electron configuration diagram for the same element in the format shown. The drawing below is only to show format and has no information about the correct structure which you need to draw. (2 pts)

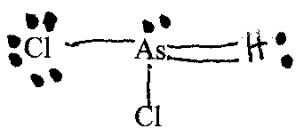


2. Give the electron configuration for the ion **Mg⁺²** (in the format 1s², 2s², 2p²...) (4 pts)

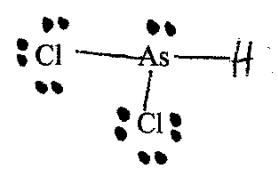
3. Give the Lewis Dot Symbol for the element **Mg** (4 pts)

4. a. For the formula **AsCl₂H** give the number of total valence electrons. Show work. (3 pts)

- b. Given the two structures drawn for the formula given above, circle the letter under the correct Lewis Dot structure (3 pts)



(a)



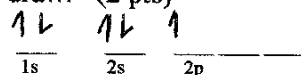
(b)

Extra Credit: Give the dipole moment arrow (+---->) for the bond shown below. (3 pts)



Name _____ Name _____
 (print name) (sign name) (I can't read some of your handwriting.)

1. a. Give the electron configuration of **Cl** in the format ($1s^2, 2s^2, \dots$ etc.) (2 pts)
- b. Give the valence electron configuration of the same element in the same format. (2 pts)
- c. Give the valence electron configuration diagram for the same element in the format shown. The drawing below is only to show format and has no information about the correct structure which you need to draw. (2 pts)

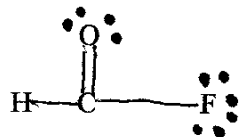


2. Give the electron configuration for the ion **Se²⁻** (in the format $1s^2, 2s^2, 2p^2, \dots$) (4 pts)

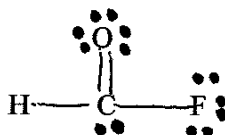
3. Give the Lewis Dot Symbol for the element **Ge** (4 pts)

- 4a. For the formula **HCOF** give the number of total valence electrons. Show work. (3 pts)

- b. Given the two structures drawn for the formula given above, circle the letter under the correct Lewis Dot structure (3 pts)



(a)



(b)

Extra Credit: Give the dipole moment arrow (+---->) for the bond shown below. (3 pts)



C ----- Br