

**TABLE 3.4 Some Common Polyatomic Ions**

Name	Formula	Name	Formula
Acetate *	$\text{C}_2\text{H}_3\text{O}_2^-$	Hypochlorite	$\text{ClO}^-$
Carbonate *	$\text{CO}_3^{2-}$	Chlorite	$\text{ClO}_2^-$
Hydrogen carbonate (or bicarbonate)	$\text{HCO}_3^-$	Chlorate	$\text{ClO}_3^-$
Hydroxide *	$\text{OH}^-$	Perchlorate	$\text{ClO}_4^-$
Nitrite	$\text{NO}_2^-$	Permanganate	$\text{MnO}_4^-$
Nitrate *	$\text{NO}_3^-$	Sulfite	$\text{SO}_3^{2-}$
Chromate	$\text{CrO}_4^{2-}$	Hydrogen sulfite (or bisulfite)	$\text{HSO}_3^-$
Dichromate	$\text{Cr}_2\text{O}_7^{2-}$	Sulfate *	$\text{SO}_4^{2-}$
Phosphate *	$\text{PO}_4^{3-}$	Hydrogen sulfate (or bisulfate)	$\text{HSO}_4^-$
Hydrogen phosphate	$\text{HPO}_4^{2-}$	Cyanide	$\text{CN}^-$
Dihydrogen phosphate	$\text{H}_2\text{PO}_4^-$	Peroxide	$\text{O}_2^{2-}$
Ammonium *	$\text{NH}_4^+$	<b>Memorize these with *</b>	