

BA = bad attempt NA = not attempted
 Please show work on all questions for full credit & partial credit. (20 total pts) NW = no work

1. Match the following with the definition or example below. (5 pts)

solvent 4 element 2 saturated solution 5 heterogeneous mixture 3 solute 1

- (1) The substance which you have less of in a solution.
 (2) N
 (3) wet sand
 (4) The substance which you have more of in a solution.
 (5) A solution in which you dissolved Li OH into water but still have some undissolved Li OH because you have added too much Li OH.

each blank 1 pt

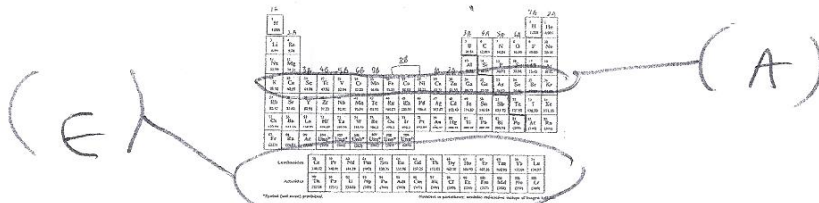
2. Give the element symbol for 2 elements which are diatomic molecules in its natural state (4 pts, 2 pts each)

-2 for one

HON + halogens
 H_2, O_2, N_2
 F_2, I_2, Br_2, Cl_2

3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts each)

(A) period (B) group (C) main group (representative elements) (D) transition metal elements (E) lanthanide/actinide



4. a. Is the element P a [(metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)

b. What is the atomic number? 15 c. What is the atomic mass? 30.97

b. How many protons? 15 How many total electrons (for a neutral atom)? 15

c. What is the group number? 5A d. How many valence electrons? 5

5. Comparing individual atoms of the elements listed, which element is bigger? [(N) or (Bi)] (circle one)

Extra Credit Question (4 pts, 2 pts each number)

bigger down group

1. Name the binary compound shown. K Cl given the name of the elements (potassium, chlorine) potassium chloride

potassium chloride

rewrite

2. Is the following compound [(ionic) or (covalent)] (circle one). Na_2S

elements opposite side table

Name Key Name _____

Sign _____ Print (bc I can't read your signatures)

~~BA = bad attempt NA = not attempt~~ green
Please show work on all questions for full credit & partial credit. (20 total pts)

1. Match the following with the definition or example below. (5 pts) NW = NO WORK

heterogeneous mixture 2 solute 5 solvent 3 element 1 saturated solution 4

(5) The substance which you have less of in a solution. ask blank - 1 pt

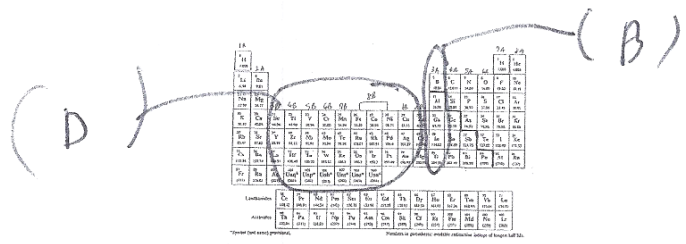
- (1) N
- (2) sand with iron filings
- (3) The substance which you have more of in a solution.
- (4) A solution in which you dissolved KCl into water with some undissolved solid KCl salt still in the bottom of the beaker. NON + halogen

2. Give the element symbol for 2 elements which are diatomic molecules in its natural state _____ (4 pts, 2 pts each) -2 for one

H₂ O₂ N₂
F₂ Cl₂ Br₂ I₂

3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts each)

(A) period (B) group (C) main group (representative elements) (D) transition metal elements (E) lanthanide/actinide



4. a. Is the element Cl a [(metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)

- b. What is the atomic number? 17 c. What is the atomic mass? 35.5
- b. How many protons? 17 How many total electrons (for a neutral atom)? 17
- c. What is the group number? 7A d. How many valence electrons? 7

5. Comparing individual atoms of the elements listed, which element is bigger? [(Cs) or (Li)] (circle one)

Extra Credit Question (4 pts, 2 pts each number)

bigger down group

1. Name the binary compound shown. AgBr given the name of the elements (silver, bromine) ^{+ide}
silver bromide

rewrite -2

2. Is the following compound [(ionic) or (covalent)] (circle one). PCl₃

close in periodic table

BA = bad attempt, NA = not attempted, NW = no work
 Please show work on all questions for full credit & partial credit. (20 total pts)

1. Match the following with the definition or example below. (5 pts)

homogeneous mixture 3 solute 1 solvent 4 element 2 saturated solution 5

- (1) The substance which you have less of in a solution.
- (2) N (nitrogen)
- (3) Coffee with water, sugar and milk well stirred.
- (4) The substance which you have more of in a solution.
- (5) A solution in which you dissolved KCl into water with some undissolved solid KCl salt still in the bottom of the beaker.

each blank 1 pt

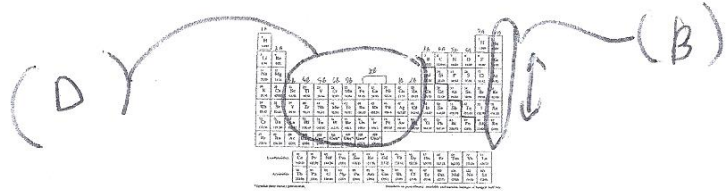
2. Give the element symbol for 2 elements which are diatomic molecules in its natural state (4 pts, 2 pts each)

F₂, Cl₂, Br₂, I₂ *halogens + H₂O, N₂*

-2 for each

3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts each)

- (A) period (B) group (C) main group (representative elements) (D) transition metal elements (E) lanthanide/actinide



4. a. Is the element Ra a [(metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)

- b. What is the atomic number? 88 c. What is the atomic mass? 226
- b. How many protons? 88 How many total electrons (for a neutral atom)? 88
- c. What is the group number? 2A d. How many valence electrons? 2

5. Comparing individual atoms of the elements listed, which element is bigger? [(Cs) or (Li)] (circle one)

Extra Credit Question (4 pts, 2 pts each number) *bigger down group*

1. Name the binary compound shown. Ba F₂ given the name of the elements (barium, fluorine)

barium fluoride *rewrote above*

2. Is the following compound [(ionic) or (covalent)] (circle one). K₂O

elements opposite side of periodic table

Name Key Name _____

Sign _____ Print (bc I can't read your signatures)

BA = bad attempt NA = not attempted green
NW = no work

Please **show work on all questions** for **full credit & partial credit**. (20 total pts)

1. Match the following with the definition or example below. (5 pts)

homogeneous mixture 1 solute 4 solvent 2 element 5 saturated solution 3

- (1) Coffee with water, sugar and milk well stirred.
- (2) The substance which you have more of in a solution.
- (3) A solution in which you dissolved KCl into water with some undissolved solid KCl salt still in the bottom of the beaker.
- (4) The substance which you have less of in a solution.
- (5) N (nitrogen)

each blank - 1 pt

2. Give the element symbol for **2 elements** which are diatomic molecules in its natural state (4 pts, 2 pts each)

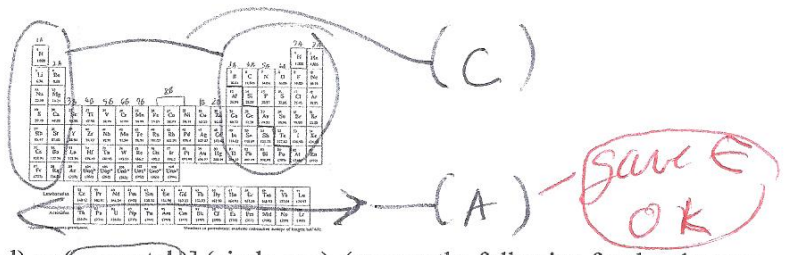
halogens + H₂

H₂, O₂, N₂
F₂, Cl₂
Br₂, I₂

-3 for one

3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts, 2 pts each)

- (A) period
- (B) group
- (C) main group (representative elements)
- (D) transition metal elements
- (E) lanthanide/actinide



4. a. Is the element Ne a [(metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)

- b. What is the atomic number? 10 c. What is the atomic mass? 20.18
- b. How many protons? 10 How many total electrons (for a neutral atom)? 10
- c. What is the group number? 8A d. How many valence electrons? 8

5. Comparing individual atoms of the elements listed, which element is bigger? [(Li) or (F)] (circle one)

Extra Credit Question (4 pts, 2 pts each number)

smaller across period
rewrote
-2

1. Name the binary compound shown. Ca I₂ given the name of the elements (calcium, iodine)^{side}

Calcium iodide

2. Is the following compound [(ionic) or (covalent)] (circle one). S O₂

elements opposite side periodic table

Name _____ Name _____

Sign _____ Print (bc I can't read your signatures)

Please **show work on all questions** for full credit & partial credit. (20 total pts)

1. Match the following with the definition or example below. (5 pts)

solvent _____ element _____ saturated solution _____ heterogeneous mixture _____ solute _____

- (1) The substance which you have less of in a solution.
- (2) N
- (3) wet sand
- (4) The substance which you have more of in a solution.
- (5) A solution in which you dissolved Li OH into water but still have some undissolved Li OH because you have added too much Li OH.

2. Give the element symbol for **2 elements** which are diatomic molecules in its natural state _____ (4 pts, 2 pts each)

3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts each)

(A) period (B) group (C) main group (representative elements) (D) transition metal elements (E) lanthanide/actinide

4. a. Is the element **P** a [(metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)

b. What is the atomic number ? _____ c. What is the atomic mass ? _____

b. How many protons ? _____ How many total electrons (for a neutral atom)? _____

c. What is the group number ? _____ d. How many valence electrons ? _____

5. Comparing individual atoms of the elements listed, which element is bigger ? [(N) or (Bi)] (circle one)

Extra Credit Question (4 pts, 2 pts each number)

1. Name the binary compound shown. K Cl given the name of the elements (potassium, chlorine)

2. Is the following compound [(ionic) or (covalent)] (circle one). Na₂S

Name _____ Name _____

Sign _____ Print (bc I can't read your signatures)

Please **show work on all questions** for full credit & partial credit. (20 total pts)

1. Match the following with the definition or example below. (5 pts)

heterogeneous mixture _____ solute _____ solvent _____ element _____ saturated solution _____

(5) The substance which you have less of in a solution.

- (1) N
- (2) sand with iron filings
- (3) The substance which you have more of in a solution.
- (4) A solution in which you dissolved KCl into water with some undissolved solid KCl salt still in the bottom of the beaker.

2. Give the element symbol for **2 elements** which are diatomic molecules in its natural state _____ (4 pts, 2 pts each)

3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts each)

(A) period (B) group (C) main group (representative elements) (D) transition metal elements (E) lanthanide/actinide

4. a. Is the element **Cl** a [(metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)

b. What is the atomic number? _____ c. What is the atomic mass? _____

b. How many protons? _____ How many total electrons (for a neutral atom)? _____

c. What is the group number? _____ d. How many valence electrons? _____

5. Comparing individual atoms of the elements listed, which element is bigger? [(Cs) or (Li)] (circle one)

Extra Credit Question (4 pts, 2 pts each number)

1. Name the binary compound shown. Ag Br given the name of the elements (silver, bromine)

2. Is the following compound [(ionic) or (covalent)] (circle one). PCl₃

Name _____ Name _____
Sign _____ Print (bc I can't read your signatures)

Please **show work on all questions** for full credit & partial credit. (20 total pts)

1. Match the following with the definition or example below. (5 pts)

homogeneous mixture _____ solute _____ solvent _____ element _____ saturated solution _____

- (1) The substance which you have less of in a solution.
- (2) N (nitrogen)
- (3) Coffee with water, sugar and milk well stirred.
- (4) The substance which you have more of in a solution.
- (5) A solution in which you dissolved KCl into water with some undissolved solid KCl salt still in the bottom of the beaker.

2. Give the element symbol for **2 elements** which are diatomic molecules in its natural state _____ (4 pts, 2 pts each)

3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts each)

- (A) period (B) group (C) main group (representative elements) (D) transition metal elements
- (E) lanthanide/actinide

4. a. Is the element Ru a [(metal) or (nonmetal)] (circle one) (**answer the following for the element given**). (8 pts, 1 pt per blank) Ra

b. What is the atomic number? _____ c. What is the atomic mass? _____

b. How many protons? _____ How many total electrons (for a neutral atom)? _____

c. What is the group number? _____ d. How many valence electrons? _____

5. Comparing individual atoms of the elements listed, which element is bigger? [(Cs) or (Li)] (circle one)

Extra Credit Question (4 pts, 2 pts each number)

1. Name the binary compound shown. BaF₂ given the name of the elements (barium, fluorine)

2. Is the following compound [(ionic) or (covalent)] (circle one). K₂O

Name _____ Name _____

Sign _____ Print (bc I can't read your signatures)

Please **show work on all questions** for **full credit & partial credit**. (20 total pts)

1. Match the following with the definition or example below. (5 pts)

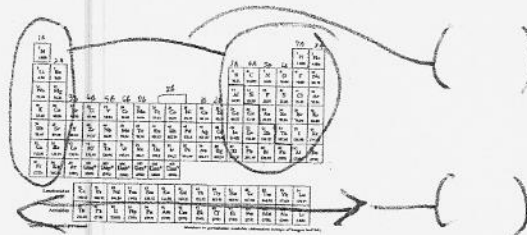
homogeneous mixture _____ solute _____ solvent _____ element _____ saturated solution _____

- (1) Coffee with water, sugar and milk well stirred.
- (2) The substance which you have more of in a solution.
- (3) A solution in which you dissolved KCl into water with some undissolved solid KCl salt still in the bottom of the beaker.
- (4) The substance which you have less of in a solution.
- (5) N (nitrogen)

2. Give the element symbol for **2 elements** which are diatomic molecules in its natural state _____ (4 pts, 2 pts each)

3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts, 2 pts each)

- (A) period (B) group (C) main group (representative elements) (D) transition metal elements
 (E) lanthanide/actinide



4. a. Is the element **Ne** a [(metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)

b. What is the atomic number ? _____ c. What is the atomic mass ? _____

b. How many protons ? _____ How many total electrons (for a neutral atom)? _____

c. What is the group number ? _____ d. How many valence electrons ? _____

5. Comparing individual atoms of the elements listed, which element is bigger ? [(Li) or (F)] (circle one)

Extra Credit Question (4 pts, 2 pts each number)

1. Name the binary compound shown. Ca I_2 given the name of the elements (calcium, iodine)

2. Is the following compound [(ionic) or (covalent)] (circle one). S O_2

