	9am			
	al Science (PSC 102) Fall 18 Dr. Hahn MWF Lyam Quiz II form A 9/7 F Exam #			
Name	Name			
Sign	Print (bc I can't read your signatures)			
Place	2 had attempt (NA = not attempted)			
Please	$\langle N \rangle \omega^{-}$			
1.	Match the following with the definition or example below. (5 pts)			
solven				
(2)	The substance which you have less of in a solution. N N N N N N N N N N N N N			
, ,	wet sand The substance which you have more of in a solution.			
. /				
(3)	A solution in which you dissolved Li OH into water but still have some undissolved Li OH because you have added too much Li OH.			
2.	Give the element symbol for 2 elements which are diatomic molecules in its natural state $\frac{1}{2}$ $\frac{1}{$			
pts eac				
3. each)	Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts			
200	(A) period (B) group (C) main group (representative elements) (D) transition metal elements (E) lanthanide/actinide			
	1			
4.	a. Is the element \underline{P} a [(metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)			
	b. What is the atomic number? $\boxed{5}$ c. What is the atomic mass? $\frac{209}{100}$			
	b. How many protons ? 15 How many total electrons (for a neutral atom)? 15			
	c. What is the group number? SA d. How many valence electrons?			
5.	Comparing individual atoms of the elements listed, which element is bigger? [(N) or Bi) (circle one)			
Extra	Credit Question (4 pts, 2 pts each number) bigger deur Group			
1.	Name the binary compound shown. K Cl given the name of the elements (potassium, chlorine)			
	potassium chloride rewrite			
2.	Is the following compound [(ionic) or (covalent)] (circle one). Na ₂ S			
((I combat's operasing & do table			

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Physical Science (PSC 102) Fall 18 Dr. Hahn MWF 11 am Quiz II form B 9/7 F Exam #
Name Name
Sign Print (bc I can't read your signatures)
Please show work on all questions for full credit & partial credit. (20 total pts)
1. Match the following with the definition or example below. (5 pts) Www however
heterogeneous mixture 2 solute 5 solvent 3 element saturated solution 4
(5) The substance which you have less of in a solution. Quilblank - 1 Rf
(1) N
(2) sand with iron filings
(3) The substance which you have more of in a solution.
(4) A solution in which you dissolved K Cl into water with some undissolved solid K Cl salt still in the
bottom of the beaker. Worthalogh
120,102
pts each) (-2 for one) F2, Cl2, B12 I2
3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts each)
(A) period (B) group (C) main group (representative elements) (D) transition metal elements (E) lanthanide/actinide
E A SA S
According to \$2 - \$7 - \$7 - \$7 - \$7 - \$7 - \$7 - \$7 -
4. a. Is the element given). (8 pts, 1 pt per blank) a [(metal) or (nonmetal)] (circle one) (answer the following for the element
b. What is the atomic number? c. What is the atomic mass? 35.5
b. How many protons ?
c. What is the group number? $\nearrow P$ d. How many valence electrons?
5. Comparing individual atoms of the elements listed, which element is bigger? [(Cs) or (Li)] (circle one)
Extra Credit Question (4 pts, 2 pts each number) (b) 3yer down gray
1. Name the binary compound shown. Ag Br given the name of the elements (silver, browning)
silver bromide (rewrote -1)
In the following a superior of the following and the following and the following a superior of the following and the following and the following and the following a superior of the following and the following a
2. Is the following compound [(ionic) or (covalent)] (oircle one). P Cl ₃
Close in poridic 14/10

ilan
Physical Science (PSC 102) Fall 18 Dr. Hahn MWF am Quiz II form A 9/7 F Exam #
Name Name
Sign Print (bc I can't read your signatures)
14 2 Dad attempt? WA = not attempted you
Please show work on all questions for full credit & partial credit. (20 total pts)
1. Match the following with the definition or example below. (5 pts)
homogeneous mixture solute solvent saturated solution
 (1) The substance which you have less of in a solution. (2) N (nitrogen) (2) N (nitrogen)
(3) Coffee with water, sugar and milk well stirred.(4) The substance which you have more of in a solution.
(5) A solution in which you dissolved K Cl into water with some undissolved solid K Cl salt still in the
1-4
bottom of the beaker.
2. Give the element symbol for <u>2 elements</u> which are diatomic molecules in its natural state $\frac{1}{\sqrt{2}}$ (4 pts, 2
pts each) Fz, Cl2, Br, I, thologen + (HON)
3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts each)
(A) period (B) group (C) main group (representative elements) (D) transition metal elements (E) lanthanide/actinide
(B) tantatando de tanta
4. a. Is the element (metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)
b. What is the atomic number?
b. How many protons ? _ A How many total electrons (for a neutral atom)? _ A
c. What is the group number? 2A d. How many valence electrons?
5. Comparing individual atoms of the elements listed, which element is bigger? [(Cs)) or (Li)] (circle one)
Extra Credit Question (4 pts, 2 pts each number) (5 i 6 g er clown group)
1. Name the binary compound shown. Ba F ₂ given the name of the elements (barium, fluorinc)
barium fluoride (centote about)
2. Is the following compound [(ionic) or (covalent)] (circle one). K ₂ O -2
2. Is the following compound [(nonic) or (covalent)] (circle one). K ₂ O -2 elements opposite side of periodic table

Physical Science (PSC 102) Fall 18 Dr. Hahn MWF 9 am Quiz II form B 9/7 F Exam #			
Name Name			
Sign Print (bc I can't read your signatures)			
BA = bad attempt (NH = Not attempted) green			
Please show work on all questions for full credit & partial credit. (20 total pts)			
1. Match the following with the definition or example below. (5 pts)			
homogeneous mixture solute solvent element saturated solution			
 (1) Coffee with water, sugar and milk well stirred. (2) The substance which you have more of in a solution. (3) A solution in which you dissolved K Cl into water with some undissolved solid K Cl salt still in the bottom of the beaker. 			
(4) The substance which you have less of in a solution.			
(5) N (nitrogen) + 02 N2			
2. Give the element symbol for <u>2 elements</u> which are diatomic molecules in its natural state [(4 pts, 2			
pts each) (halogest + Tlow) Brz, Iz			
3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts, 2 pts each)			
(A) period (B) group (C) main group (representative elements) (D) transition metal elements (E) lanthanide/actinide			
4. a. Is the element Ne a [(metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)			
b. What is the atomic number? c. What is the atomic mass? &			
b. How many protons ? 10 How many total electrons (for a neutral atom)? 10			
c. What is the group number? & A d. How many valence electrons?			
5. Comparing individual atoms of the elements listed, which element is bigger ? [(Li)) or (F)] (circle one)			
Extra Credit Question (4 pts, 2 pts each number)			
1. Name the binary compound shown. Ca I2 given the name of the elements (calcium, iodine)			
Calcium iodide rewrote			
2. Is the following compound [(ionic))or (covalent)] (circle one). S O ₂			
Clements Opposite side periodic table			

	cience (PSC 102) Fall 18 Dr. Hahn MWF 11 am Quiz II form A 9/7 F Exam #Name
Sign	Name Print (bc I can't read your signatures)
Please sho	w work on all questions for full credit & partial credit. (20 total pts)
1. Ma	atch the following with the definition or example below. (5 pts)
solvent _	element saturated solution heterogeneous mixture solute
(2) N (3) we (4) The (5) A s	ne substance which you have less of in a solution. et sand e substance which you have more of in a solution. solution in which you dissolved Li OH into water but still have some undissolved Li OH cause you have added too much Li OH.
2. Giv pts each)	we the element symbol for <u>2 elements</u> which are diatomic molecules in its natural state(4 pts,
3. Giv	ven the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts
(A) period lanthanide/a	
4. a. give	Is the element $\underline{\mathbf{P}}$ a [(metal) or (nonmetal)] (circle one) (answer the following for the element en). (8 pts, 1 pt per blank)
b.	What is the atomic number? c. What is the atomic mass?
b.	How many protons ? How many total electrons (for a neutral atom)?
c.	What is the group number? d. How many valence electrons?
5. Con	nparing individual atoms of the elements listed, which element is bigger? [(N) or (Bi)] (circle one)
Extra Cred	it Question (4 pts, 2 pts each number)
l. Nar	me the binary compound shown. K Cl given the name of the elements (potassium, chlorine)
. Is th	ne following compound [(ionic) or (covalent)] (circle one). Na ₂ S

Physica Name	al Science (PSC 102) Fall 18 Dr. Hahn MWF 11 am Quiz II form B 9/7 F Exam #Name
Sign	Print (bc I can't read your signatures)
Please	show work on all questions for full credit & partial credit. (20 total pts)
1.	Match the following with the definition or example below. (5 pts)
. \	e substance which you have less of in a solution.
(3) (4)	N sand with iron filings The substance which you have more of in a solution. A solution in which you dissolved K Cl into water with some undissolved solid K Cl salt still in the bottom of the beaker.
2. pts each	Give the element symbol for <u>2 elements</u> which are diatomic molecules in its natural state(4 pts, 2
3.	Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts each)
(A) perio	od (B) group (C) main group (representative elements) (D) transition metal elements (E) de/actinide
	a. Is the element CI a [(metal) or (nonmetal)] (circle one) (answer the following for the element (8 pts, 1 pt per blank)
1	b. What is the atomic number? c. What is the atomic mass?
1	b. How many protons ? How many total electrons (for a neutral atom)?
C	What is the group number? d. How many valence electrons?
5. (Comparing individual atoms of the elements listed, which element is bigger ? [(Cs) or (Li)] (circle one)
Extra C	redit Question (4 pts, 2 pts each number)
1. N	Name the binary compound shown. Ag Br given the name of the elements (silver, bromine)

P Cl₃

2. Is the following compound [(ionic) or (covalent)] (circle one).

17 1 17 1 17 1 17 1 1 1 1 1 1 1 1 1 1 1	02) Fall 18 Dr. Hahn MWF	9 am Quiz II form A 9/7 F Exam #
Sign		Print (bc I can't read your signatures)
	8	partial credit. (20 total pts)
	ing with the definition or exa	
homogeneous mixture _	solute sol	vent element saturated solution
(2) N (nitrogen)(3) Coffee with wate(4) The substance wh(5) A solution in whibottom of the bea	ker.	d. elution. water with some undissolved solid K Cl salt still in the
pts each)	Mildol for <u>2 elements</u> which are	e diatomic molecules in its natural state (4 pts, 2
3. Given the following	g miniaturized periodic table, fi	ill in the blank with the correct letter. (4 pts total, 2 pts each)
(A) period (B) group (C (E) lanthanide/actinide) main group (representative el	ements) (D) transition metal elements
4. a. Is the element given). (8 pts, 1 pt per bland	ent Rú a [(metal) or (nonme	tal)] (circle one) (answer the following for the element
b. What is the	atomic number? c	. What is the atomic mass?
b. How many	protons ? How many	total electrons (for a neutral atom)?
c. What is the	group number? d.	How many valence electrons ?
5. Comparing individu	al atoms of the elements listed	, which element is bigger? [(Cs) or (Li)] (circle one)
Extra Credit Question (4	pts, 2 pts each number)	
1. Name the binary c	ompound shown. Ba F ₂ g	given the name of the elements (barium, fluorine)
2. Is the following co	ompound [(ionic) or (covaler	nt)] (circle one). K ₂ O

Phys Nam	hysical Science (PSC 102) Fall 18 Dr. Hahn MWF 9 am Quiz II	form B 9/7 F Exam #
Sign		an't read your signatures)
Pleas	lease show work on all questions for full credit & partial credit	
1.	Match the following with the definition or example below. (5 pts)
homo	omogeneous mixture solute solvent	element saturated solution
(3	 (1) Coffee with water, sugar and milk well stirred. (2) The substance which you have more of in a solution. (3) A solution in which you dissolved K Cl into water with sor bottom of the beaker. (4) The substance which you have less of in a solution. (5) N (nitrogen) 	ne undissolved solid K Cl salt still in th
2. pts ea	Give the element symbol for <u>2 elements</u> which are diatomic molest each)	cules in its natural state (4 pts, 2
3.	Given the following miniaturized periodic table, fill in the blank w	with the correct letter. (4 pts, 2 pts each)
(E) lar	a) period (B) group (C) main group (representative elements) (D) train and a second se)
4.	4. a. Is the element <u>Ne</u> a [(metal) or (nonmetal)] (circle one) given). (8 pts, 1 pt per blank)	(answer the following for the element
	b. What is the atomic number? c. What is the at	tomic mass ?
	b. How many protons ? How many total electrons (for a neutral atom)?
	c. What is the group number? d. How many va	alence electrons ?
5.	5. Comparing individual atoms of the elements listed, which element	is bigger ? [(Li) or (F)] (circle one)
Extra	tra Credit Question (4 pts, 2 pts each number)	
1.	1. Name the binary compound shown. Ca I ₂ given the name	of the elements (calcium, iodine)
2.	2. Is the following compound [(ionic) or (covalent)] (circle one)	202