

Physical Science (PSC 102) Fall 18 Dr. Hahn MWF ¹¹ am Quiz II form A 9/7 F Exam # _____
 Name _____ Name _____
 Sign _____ Print (bc I can't read your signatures)

extra credit quiz redo

Please show work on all questions for **full credit & partial credit**. (20 total pts)

1. Match the following with the definition or example below. (5 pts)

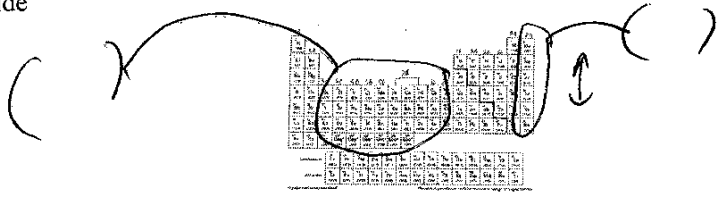
homogeneous mixture _____ solute _____ solvent _____ element _____ saturated solution _____

- (1) The substance which you have less of in a solution.
- (2) N (nitrogen)
- (3) Coffee with water, sugar and milk well stirred.
- (4) The substance which you have more of in a solution.
- (5) A solution in which you dissolved KCl into water with some undissolved solid KCl salt still in the bottom of the beaker.

2. Give the element symbol for **2 elements** which are diatomic molecules in its natural state _____ (4 pts, 2 pts each)

3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts each)

- (A) period (B) group (C) main group (representative elements) (D) transition metal elements
- (E) lanthanide/actinide



4. a. Is the element ~~Ba~~ ^{Ra} a [(metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)

- b. What is the atomic number ? _____ c. What is the atomic mass ? _____
- b. How many protons ? _____ How many total electrons (for a neutral atom)? _____
- c. What is the group number ? _____ d. How many valence electrons ? _____

5. Comparing individual atoms of the elements listed, which element is bigger ? [(Cs) or (Li)] (circle one)

Extra Credit Question (4 pts, 2 pts each number)

- 1. Name the binary compound shown. Ba F₂ given the name of the elements (barium, fluorine)
- 2. Is the following compound [(ionic) or (covalent)] (circle one). K₂O

Physical Science (PSC 102) Fall 18 Dr. Hahn MWE ¹¹ am Quiz II form B 9/7 F Exam # _____
 Name _____ Name _____
 Sign _____ Print (bc I can't read your signatures)

extra credit quiz redo

green

Please **show work on all questions** for **full credit & partial credit**. (20 total pts)

1. Match the following with the definition or example below. (5 pts)

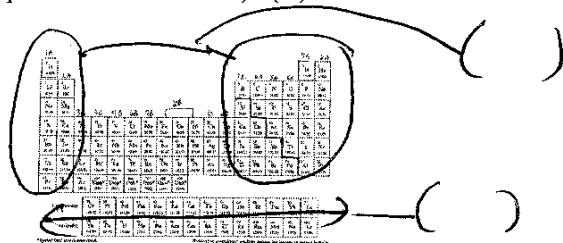
homogeneous mixture _____ solute _____ solvent _____ element _____ saturated solution _____

- (1) Coffee with water, sugar and milk well stirred.
- (2) The substance which you have more of in a solution.
- (3) A solution in which you dissolved KCl into water with some undissolved solid KCl salt still in the bottom of the beaker.
- (4) The substance which you have less of in a solution.
- (5) N (nitrogen)

2. Give the element symbol for **2 elements** which are diatomic molecules in its natural state _____ (4 pts, 2 pts each)

3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts, 2 pts each)

(A) period (B) group (C) main group (representative elements) (D) transition metal elements
 (E) lanthanide/actinide



4. a. Is the element **Ne** a [(metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)

b. What is the atomic number? _____ c. What is the atomic mass? _____

b. How many protons? _____ How many total electrons (for a neutral atom)? _____

c. What is the group number? _____ d. How many valence electrons? _____

5. Comparing individual atoms of the elements listed, which element is bigger? [(Li) or (F)] (circle one)

Extra Credit Question (4 pts, 2 pts each number)

1. Name the binary compound shown. Ca I₂ given the name of the elements (calcium, iodine)

2. Is the following compound [(ionic) or (covalent)] (circle one). S O₂

Physical Science (PSC 102) Fall 18 Dr. Hahn MWF 11 am Quiz II form A 9/7 F Exam # _____
 Name _____ Name _____
 Sign _____ Print (bc I can't read your signatures)

extra credit quiz redo

Please **show work on all questions** for **full credit & partial credit**. (20 total pts)

1. Match the following with the definition or example below. (5 pts)

solvent _____ element _____ saturated solution _____ heterogeneous mixture _____ solute _____

- (1) The substance which you have less of in a solution.
- (2) N
- (3) wet sand
- (4) The substance which you have more of in a solution.
- (5) A solution in which you dissolved Li OH into water but still have some undissolved Li OH because you have added too much Li OH.

2. Give the element symbol for **2 elements** which are diatomic molecules in its natural state _____ (4 pts, 2 pts each)

3. Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts each)

(A) period (B) group (C) main group (representative elements) (D) transition metal elements (E) lanthanide/actinide

4. a. Is the element **P** a [(metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)

b. What is the atomic number ? _____ c. What is the atomic mass ? _____

b. How many protons ? _____ How many total electrons (for a neutral atom)? _____

c. What is the group number ? _____ d. How many valence electrons ? _____

5. Comparing individual atoms of the elements listed, which element is bigger ? [(N) or (Bi)] (circle one)

Extra Credit Question (4 pts, 2 pts each number)

1. Name the binary compound shown. K Cl given the name of the elements (potassium, chlorine)

2. Is the following compound [(ionic) or (covalent)] (circle one). Na₂S

9am

Physical Science (PSC 102) Fall 18 Dr. Hahn MWF 11 am Quiz II form B 9/7 F Exam # _____

Name _____ Name _____

Sign _____ Print (bc I can't read your signatures)

extra credit quiz redo

green

Please **show work on all questions** for **full credit & partial credit**. (20 total pts)

1. Match the following with the definition or example below. (5 pts)

solvent _____ element _____ saturated solution _____ heterogeneous mixture _____ solute _____

(5) The substance which you have less of in a solution.

(1) N

(2) Sand with iron filings

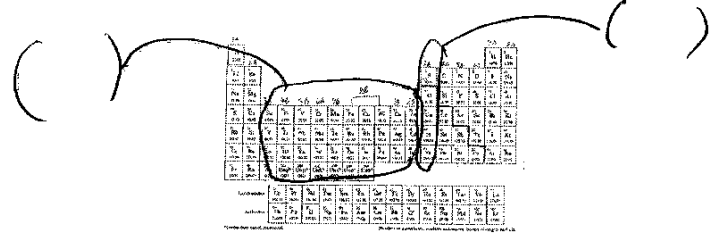
(3) The substance which you have more of in a solution.

(4) A solution in which you dissolved KCl into water but with some undissolved KCl salt still in the bottom of the beaker.

(6) Give the element symbol for **2 elements** which are diatomic molecules in its natural state _____ (4 pts, 2 pts each)

(7) Given the following miniaturized periodic table, fill in the blank with the correct letter. (4 pts total, 2 pts each)

(A) period (B) group (C) main group (representative elements) (D) transition metal elements (E) lanthanide/actinide



(8) a. Is the element **Cl** a [(metal) or (nonmetal)] (circle one) (answer the following for the element given). (8 pts, 1 pt per blank)

b. What is the atomic number? _____ c. What is the atomic mass? _____

b. How many protons? _____ How many total electrons (for a neutral atom)? _____

c. What is the group number? _____ d. How many valence electrons? _____

(9) Comparing individual atoms of the elements listed, which element is bigger? [(Cs) or (Li)] (circle one)

Extra Credit Question (4 pts, 2 pts each number)

1. Name the binary compound shown. Ag Br given the name of the elements (silver, bromine)

2. Is the following compound [(ionic) or (covalent)] (circle one). PCl₃