Quiz III	General Chemistry	y I Lecture Spring	g 15 Dr. Hahn	20 pts 2/1	2 R 8:30a	m form A	quiz#
Name	Ken		Name				
(print nar	me)		(sign nar	ne)			
1 0	. 1 4 6 11 .	1. 17.1				.11	(4 -4-)
1. 0	ircle the following	compounds which	are ionic You	may circle c	one, many, a	iii or none.	(4 pts)
MgBr <sub>2</sub>	SO <sub>2</sub> H <sub>2</sub> O	Na <sub>2</sub> SO <sub>4</sub> )					
	That are 2 diatomic the element symbols $F_{2}$ , $Cl_{2}$ ,	) any 2 diatomic e	elements. (4	pts)		atomic)? Na	ame (or give
3. W	rite out the correct	ionic formula for	the following.	Show work	(6 pts, wo	ork 3 pts)	
Between	the elements: Al	and P III	$A \rightarrow +3$	J+ Me	voy I	vet Vet	9-8=-5 \
	AS	P	100	Long	3)(	owest Vatus	
4 - F-1	7	et each)	N	W-1	61	1- (6 -1-)	
Na <sub>2</sub> SO <sub>4</sub>	or the following for $N_{\alpha}$	22,999	Imal	5-7	32.01	3/mal	(0-)
2(2	2,99)+	32.07 +	4(16	.00) =	142,00	5 9/m n	16,003/ml
1pt-h	latet	1st	Ipt 0	lot		Jingo	
	dit (2 pts) Given tule, how many mo	he following mole	cule and the gi	ven formula	mass. If you		55 grams of
H Cl (FW	V = 36.46  g/mol		(1)		<i>σ</i>	2 10	
37	L,55g H(	1 × 1/36	of Help.	= 0,	892.	f hal	HU
		Det .	The state of the s	)			

k n	ture Spring 15 Dr. Hahn 20 pts 2/12 R	4 ellow
Name (print name)	Name (sign_name)	
(6)	(orgin munico)	
1 Circle the fellowing a survey	and a little and the SZ	
1. Choic the following compo	ounds which are covalent. You may circle of	
Man. (go) (Ho) M	so4 not lowest (VW	
$MgBr_2$ $(SO_2)$ $(H_2O)$ $Na_2$	SO4 (Northo In )	1 50 11 1 60
		- / 3 \
2. What are 2 diatomic eleme	nts (elements which in the most stable state diatomic elements. (4 pts)	e is a diatomic)? Name (or give
	, = <i>/</i>	a/ (Frat)
12 /2 Brz	Iz Hz Oz	2
		(FR-1)
3. Write out the correct ionic	formula for the following. Show work (6	
Between the elements: Ca and	CI - Grazetti A	7-8=-1 (Charge pt)
(	•	ener
gre	up IA +2	r (Caty Cet)
$((17)^{2}(+2) + ($	#CV(-1) = Zero	
( Cas	to Call	La Co
4 For the following formula	give the formula mass (or molar mass). Sh	
AI (NO3)3 AL-LO, 48	S/mel N-14.019/mol (	)-16,00 g/mel
1 3 9		
r i nl	4 1	
	$\sim$	
1(26,	70) 2 (1, 21) 1 0 1	
1/20,	18) + 3(1401) + 9(	(16,00) = 2/3,019/mol
extra Credit (2 pts)—Given the following the molecule, how many moles d	o you have? Show work	s. If you have 543:2 grams
	(ext	raster-tot
$I_2O \text{ (FW = 18.02 g/mol)}$		
543,7 a HOX	maltro 5 = 30,14	mol to
J K	1800 a HOU 10,11	
	102/20	

Quiz III General Chemistry I Lecture Spring 15 Dr. Hahn 20 pts 2/12 R 9:55 am form A quiz #
Name Name not lowest vatro 1
(print name) (sign name) WWF No charge -3)
1. Circle the following which is an empirical formula (4 pts)
C2H2 CH) NH2) N2H4 ) LOWest whole # Vatur's
2. What are 2 diatomic elements (elements which in the most stable state is a diatomic)? Name (or give)
the element symbol) any 2 diatomic elements. (4 pts)
F2 Cl2 Br2 IZ H2 O2 N2( [Mar))
3. Write out the correct ionic formula for the following. Show work (6 pts, work 3 pts)
Capacta tal
Between the elements: Rb and N
group IA +1
(#Rb)(+1) + (#N)(-3) = Zero) (Rb3 N)(NW-10)
1 Rho N Charles
The second secon
4. Give the name or the formula and charge (for the polyatomic ions) or number for the following (6 pts, 3 pts each)
The state of the s
Sulfate 504 (-2) HCI hy dro Chloric acid
- 1 CO 4 1D
Extra Credit (2 pts)— Given the following molecule and the given formula mass. If you have 160.2 grams
of the molecule, how many moles do you have? Show work.
Na <sub>3</sub> PO <sub>4</sub> (FW = 163.94 g/mol)
160,29 x mel Naz PO45 = 0,9772 mel Naz PO4 163,949 & 1/2 RD
Naz PO4 165,749 4 12 20
Naz P04 ( Naz P04
(-12pt) (-Int)

Tame	Ken	Name	pink
print name)	0	(sign name)	
1 Circle the	fallowing which is a	-landar Consultation (A at )	
1. Circle the	ير -	iolecular formula (4 pts) (NOTEMPINI	C = 1\
$^{2}H_{2}$	CH NH <sub>2</sub> (N	2H4)	Call)
· ·			
2. What are the eleme	2 diatomic elements (el ent symbol) any 2 diator	ements which in the most stable stance elements. (4 pts)	ite is a diatomic)? Name (or give
_	1	` • ′	1 ( that
Fz C	2 15/2 -	-2 H2 O2	102 FO-1
3. Write out	the correct ionic formu	la for the following. Show work	(6 pts, work 3 pts)
etween the eler	nents: Sr and N	~ IA -> 5-8=	= 3 (Chargo loA)
	1		- 3 Charge last
	JIA-	1 ,	NO NO
(40)C	+2) + (#N)(- 1 2	2) 70	(NW-IRT)
	79 T (#N/C	5)= tenb	
\ 2	2 (	SKAN) /-	N2 N2 NWH
	·	23 2	3 hoch arge
4. Give the n	ame or the formula and	charge (for the polyatomic ions) o	r number for the following (6 -2
pts, 3 pts	each not mout u	3) (Large-tet	notlower
	-3 A		rate - lot
(as number pref	ix) perta	phosphate P04	
	<i>y</i> • , • • •		
tra Credit (2 pt	s) Given the following	molecule and the given formula ma	ass. If you have 0.2334 grams
the molecule, h	ow many moles do you	have? Show work.	
$HCO_3$ (FW =	84.01 g/mol)		
() 2334	ia ha	NaHCO3n	0 ^ -3
NIOH	7 X -840	16 2,1	18 X 10 mal
Nam	0710	A L	78 × 10 <sup>-3</sup> mal HCO3
(	Na H	13 / Wa	VI CO3

		Name	
(print name)		Name (sign name)	
1. Circle the fo	ollowing compounds v	which are ionic You may circle one, many, all or none. (	4 pts
MgBr <sub>2</sub> SO <sub>2</sub>	H <sub>2</sub> O Na <sub>2</sub> SO <sub>4</sub>		
2. What are 2 the element	diatomic elements (ele symbols) any 2 diato	ements which in the most stable state is a diatomic)? Namomic elements. (4 pts)	ie (oi
3. Write out th	e correct ionic formul	a for the following. Show work (6 pts, work 3 pts)	
Between the eleme	ents: Al and P		
	owing formula give th	ne formula mass (or molar mass). Show work. (6 pts)	
4. For the follo		to formula mass (or motal mass). Show work, (o pts)	
4. For the followant Na <sub>2</sub> SO <sub>4</sub>		to formula mass (of molar mass). Show work. (o pis)	
		to formula mass (of molar mass). Show work. (o pis)	
		to formula mass (of molar mass). Show work. (o pis)	
		to formula mass (of molar mass). Show work. (o pis)	
Na <sub>2</sub> SO <sub>4</sub> Extra Credit (2 pts)		molecule and the given formula mass. If you have 32.55	grai

Name	Name
(print name)	Name (sign name)
1. Circle the following compoun	nds which are covalent. You may circle one, many, all or none.
MgBr <sub>2</sub> SO <sub>2</sub> H <sub>2</sub> O Na <sub>2</sub> SO	O <sub>4</sub>
2. What are 2 diatomic element the element symbol) any 2 di	is (elements which in the most stable state is a diatomic)? Name (intomic elements. (4 pts)
3. Write out the correct ionic fo	ormula for the following. Show work (6 pts, work 3 pts)
Between the elements: Ca and	Cl
	•
4. For the following formula, gi	ive the formula mass (or molar mass). Show work. (6 pts)
Al (NO <sub>3</sub> ) <sub>3</sub>	•
( 3/2	
Extra Credit (2 pts) Given the follow of the molecule, how many moles do	wing molecule and the given formula mass. If you have 543.2 go you have? Show work.

Name	Name
(print n	Name (sign name)
1.	Circle the following which is an empirical formula (4 pts)
C <sub>2</sub> H <sub>2</sub>	CH $NH_2$ $N_2H_4$
2.	What are 2 diatomic elements (elements which in the most stable state is a diatomic)? Name (or give the element symbol) any 2 diatomic elements. (4 pts)
	Write out the correct ionic formula for the following. Show work (6 pts, work 3 pts) on the elements: Rb and N
4.	Give the name or the formula and charge (for the polyatomic ions) or number for the following. (6 pts, 3 pts each)
Sulfate	H C1
Extra C of the m	Gredit (2 pts) Given the following molecule and the given formula mass. If you have 160.2 grams nolecule, how many moles do you have? Show work.
Na <sub>3</sub> PO	$O_4(FW = 163.94 \text{ g/mol})$

Quiz II	I General Chemi	stry I Lecture	Spring 15 D	r. Hahn 20 pts	2/12 R 9:55 ar	m form B quiz#	
Name			N	ame			
(print i	name)		(8	sign name)			
1.	Circle the follow	ing which is a n	nolecular form	nula (4 pts)			
$C_2H_2$	СН	NH <sub>2</sub> N	<sub>2</sub> H <sub>4</sub>				
2.	What are 2 diator the element symb				ole state is a dia	tomic)? Name (or	give
	Write out the coren the elements:		ıla for the fol	lowing. Show w	ork (6 pts, wo	ork 3 pts)	
4.		the formula and	d charge (for	the polyatomic ic	ons) or number	for the following.	(6
	pts, 3 pts each)						
5 (as n	umber prefix)		phospha	te	***************************************		
Extra (of the 1	Credit (2 pts) Giv nolecule, how ma	en the following	g molecule and have? Sho	nd the given form wwork.	ula mass. If yo	ou have 0.2334 gra	ams
NaHC	$O_3$ (FW = 84.01)	g/mol)					