

Name Key (print name) NA = not attempt Name NW = no work (sign name)

Please show all work for full credit and to get partial credit. "Confidence Booster"

1. In the following number, how many significant figures are in the number given? Circle either "ambiguous" or "not ambiguous". If ambiguous give at least one possible correct number of significant figures. (6 pts)

0.08890 [(ambiguous) or (not ambiguous)] (circle one) # significant figures 4

2. Fill in the blanks. (6 pts)

a. 1 kiloliter = 1000 liters 1/1000 3pt

3. Convert 1.34 cups into liters (4 cups = 1 quart, 1.06 quart = 1 liter) (show work) (8 pts)

$$1.34 \text{ cups} \times \frac{1 \text{ quart}}{4 \text{ cups}} \times \frac{1 \text{ liter}}{1.06 \text{ quart}} = 0.316 \text{ liters}$$

Handwritten notes: 2pt, 2pt, 2pt, 2pt, 3pt, 3pt

Upside down or wrong 1 pt NW = 4pt math 1/2 pt
attempt -3 extra step -1 pt

Extra Credit: 3 pts

1. Circle the following which is an element (not a compound). (You may circle all or none.) (1/2 pts each, 1.5 pt)

O₂ CO₂ C 1/2 pt each

2. If the mass of a substance is 2.45 grams and the volume occupied by that substance is 2.8 mL, what is the density? Show work. (density = mass / volume) (1.5 pt)

$$\text{mass} = 2.45 \text{ g}, \text{ Volume} = 2.8 \text{ mL}$$

$$\text{density} = \frac{2.45 \text{ g}}{2.8 \text{ mL}} = 0.875 \frac{\text{g}}{\text{mL}}$$

NW -1 math 1/2 pt = 0.88 g/mL 2 sig fig

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108.20 [(ambiguous) or (not ambiguous)] (circle one) # significant figures 5

2. Fill in the blanks. (6 pts)

a. 100 centigrams = 1 gram

3. Convert 4.23 pounds into kilograms. (453.6 grams = 1 pound) (8 pts)

$$4.23 \text{ pounds} \times \frac{453.6 \text{ g}}{1 \text{ pound}} \times \frac{1 \text{ kg}}{1000 \text{ g}} = 1.92 \text{ kg}$$

2pt

2pt

2pt

2pt

NW = -4pt

attempt -3 pt

math -1/2 pt

extra step -1 pt

up side down -1 pt or half wrong

Extra Credit: 3 pts

1. Circle the following which is a compound (not an element). (You may circle all or none.) (1/2 pts each, 1.5 pt)

Fe F₂ Fe₂O₃

1/2 pt each

2. If the mass of a substance is 1.7889 grams and the volume occupied by that substance is 1.3 mL, what is the density? Show work. (density = mass / volume) (1.5 pt)

mass = 1.7889g volume = 1.3ml

$$\text{density} = \frac{1.7889 \text{ g}}{1.3 \text{ ml}} = 1.4 \text{ g/ml}$$

math -1/2 pt

NW -1pt

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1. In the following number, how many significant figures are in the number given? Circle either "ambiguous" or "not ambiguous". If ambiguous give at least one possible correct number of significant figures. (6 pts)

1.033000 [(ambiguous) or (not ambiguous)] (circle one) # significant figures 7

2. Fill in the blanks. (6 pts)

a. 1000 milligrams = 1 gram

$\frac{1}{1000}$ - 3pt

3. Convert 1.234 feet into centimeters (12 inches = 1 foot, 2.54 cm = 1 inch) (8pt)

$$1.234 \text{ feet} \times \frac{12 \text{ inches}}{1 \text{ foot}} \times \frac{2.54 \text{ cm}}{1 \text{ inch}} = 37.61 \text{ cm}$$

(2pt) (2pt) (2pt) (2pt)

NW - 4pt

Attempt - 3pt

math - 1pt

extra step - 1pt

up side down - 1pt or half wrong

Extra Credit: 3 pts

1. Circle the following which is an compound (not an element). (You may circle all or none.) (1/2 pts each, 1.5 pt)

NaCl Na Cl₂

1/2 pt each

NW - 1pt

math - 1/2 pt

2. If the mass of a substance is 7.99 grams and the volume occupied by that substance is 10.3 mL, what is the density? Show work. (density = mass / volume) (1.5 pt)

mass = 7.99g Volume = 10.3ml

$$\text{density} = \frac{7.99 \text{ g}}{10.3 \text{ ml}} = 0.776 \text{ g/ml}$$

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32500 [(ambiguous) or (not ambiguous)] (circle one) # significant figures 3 or 5
 (3 pt)

2. Fill in the blanks. (6 pts)

a. 1000 gram = 1 kilogram
 (1000 - 3 pt)

3. Convert 78.9 pints into milliliters. (~~2 cups = 1 pint~~, 2 pints = 1 quart, 1.06 quart = 1 liter) (8 pt)

$$78.9 \text{ pint} \times \frac{1 \text{ quart}}{2 \text{ pints}} \times \frac{1 \text{ liter}}{1.06 \text{ quart}} \times \frac{1000 \text{ ml}}{1 \text{ l}} =$$

(2 pt) (2 pt) (2 pt) (2 pt)

37216.98 calculator

$$3.72 \times 10^4 \text{ ml}$$

math - 1 pt

extra step - 1 pt

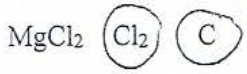
Upside down - 1 pt

Extra Credit: 3 pts

NW = 4 pt

attempt - 3 pt

1. Circle the following which is an element (not a compound). (You may circle all or none.) (1/2 pts each, 1.5 pt)



1/2 pt each

NW = -1 pt

math - 1/2 pt

2. If the mass of a substance is 1.2754 grams and the volume occupied by that substance is 6.8 mL, what is the density? Show work. (density = mass/volume) (1.5 pt)

1.2754 g = mass 6.8 ml = volume

$$\text{density} = \frac{1.2754 \text{ g}}{6.8 \text{ ml}} = 0.18756 \rightarrow 0.19$$

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Extra Credit: 3 pts

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2. If the mass of a substance is 2.45 grams and the volume occupied by that substance is 2.8 mL, what is the density? Show work. (density = mass / volume) (1.5 pt)

Quiz I General Chemistry I Lecture Spring 15 Dr. Hahn 20 pts 1/20 T 8:30am form B quiz # 5-12

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Extra Credit: 3 pts

1. Circle the following which is an element (not a compound). (You may circle all or none.) (1/2 pts each, 1.5 pt)

MgCl₂ Cl₂ C

2. If the mass of a substance is 1.2754 grams and the volume occupied by that substance is 6.8 mL, what is the density? Show work. (density = mass / volume) (1.5 pt)