HW 14.7:   Le Châtelier's Principle: If a stress is applied to a reaction mixture *at equilibrium*, net reaction occurs in the direction that relieves the stress. (equilibrium shifts to undo stress applied to system)

2 A (g) + B(g) →3 C(g) + 2 D(g)   ΔH = negative (circle one parenthesis under
 which direction will equilibrium shift   each letter)
  hint:ΔH negative is exothermic = (+ heat, heat is product)

(a)   add A   rxn goes [(→) or (←)]

(b) remove B   rxn goes [(→) or (←)]

(c) add C   rxn goes [(→) or (←)]

(d) remove C   rxn goes [(→) or (←)]

(e)   higher T   rxn goes [(→) or (←)] end Test 3 (c) & below
(higher T = add heat)
(f) higher P   rxn goes [(→) or (←)]
(assume all reactants & products in reaction are gases) (moves to fewer moles gas)

(g)   higher V   rxn goes [(→) or (←)]
(assume all reactant & products in reaction are gases)   (moves to fewer moles gas)